

COMMON ELEMENTS and COMPOUNDS**Atomic, Ionic and Equivalent Weights**

Common Name	Chemical Symbol	Common oxidation number or ionic charge	Atomic molecular or ionic mass	Equivalent Weight (*)
Aluminium	Al	3	27.0	9.0
Boron	B	3	10.8	3.5
Calcium	Ca	2	40.1	20.0
Calcium oxide	CaO		56.1	28.0
Carbon	C	4	12.0	3.0
bicarbonate	HCO ₃	-1	61.0	61.0
carbonate	CO ₃	-2	60.0	30.0
Chloride	Cl	-1	35.5	35.5
Chromium	Cr	6 (or 3)	52.0	8.7
dichromate	Cr ₂ O ₇	2	216.0	36.0
Gypsum	CaSO ₄ .2H ₂ O		172.2	86.1
Hydrogen	H	1	1.0	1.0
Iron(III)(ferric)	Fe	3	55.8	18.6
Iron(II)(ferrous)	Fe	2	55.8	27.9
Magnesium	Mg	2	24.3	12.2
oxide	MgO		40.3	20.1
Manganese (IV)(manganic)	Mn	4	54.9	13.7
Manganese (II)(manganous)	Mn	2	54.9	27.5
Nitrogen	N	3	14.0	4.7
nitrate	NO ₃	-1	62.0	62.0
Oxygen	O	2	16.0	8.0
Phosphorus	P	3 (or 5)	31.0	10.3
dihydrogen phosphate	H ₂ PO ₄	-1	97.0	97.0
hydrogen phosphate	HPO ₄	-2	96.0	48.0
phosphate	PO ₄	-3	95.0	31.7
Potassium	K	1	39.1	39.1
oxide	K ₂ O		94.2	47.1
Silicon	Si	4	28.1	7.0
Sodium	Na	1	23.0	23.0
hydroxide	NaOH		40.0	40.0
Sulphur	S	2	32.1	16.0
hydrogen sulphate	HSO ₄	-1	97.1	97.1
sulphate	SO ₄	-2	96.1	48.0
sulphide	S	-2	32.1	16.0

* Equivalent wt of an element = atomic mass divided by valency